

First reversible protonation of the all-nitrogen 1-aryl pentazole ring

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Abstract—At $-40\text{ }^{\circ}\text{C}$ in chlorosulfonic acid 1-aryl pentazoles were protonated at N-3. The N-3 atom showed a ^{15}N NMR shielding shift of -60 to -62 ppm. The $-\text{N}_5$ ring is the highest member of the azole series.

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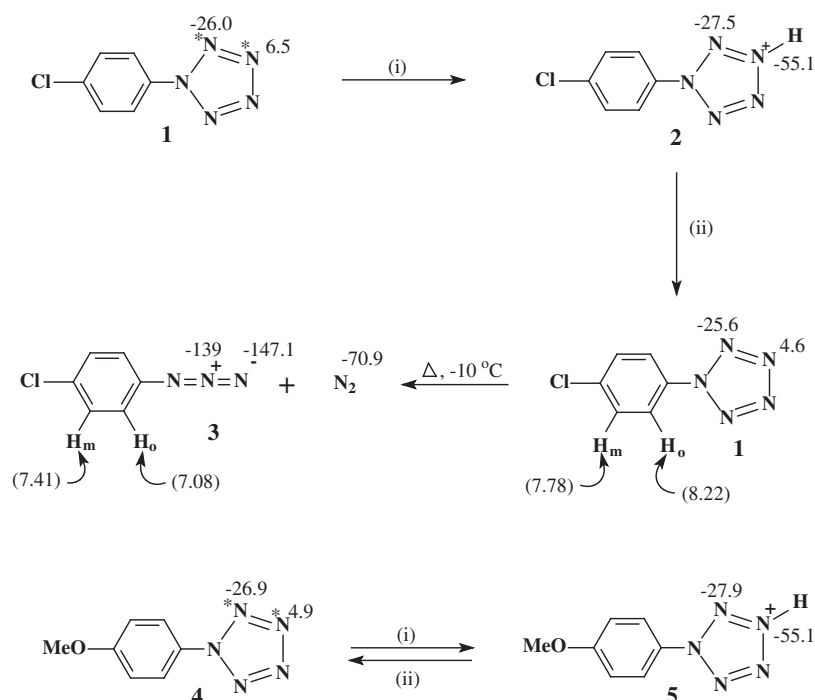
The existence of an unstable all-nitrogen five-membered ring, ArN_5 , the aryl pentazole system was first proved by Huisgen and Ugi¹ who isolated examples at low temperatures in 1956. At normal ambient temperatures the ring degrades rapidly with N_2 evolution giving aryl azides $\text{Ar}-\text{N}_3$. Since the discovery of the $\text{Ar}-\text{N}_5$ ring there has been much theoretical interest in its properties.² Little is known about its chemical behaviour other than the N_2 evolution. Attempted reductions and oxidations have all resulted in destruction of the $\text{Ar}-\text{N}_5$ ring.²⁻⁴ One chemical property that has attracted particular theoretical attention is its basicity and its protonation site.⁵⁻⁸ A very weakly basic $\text{p}K_{\text{a}}$ of ca. -8.9 ($\text{p}K_{\text{a}}$ of the conjugate acid) has been theoretically estimated^{7,8} for $\text{Me}-\text{N}_5$ and a protonation site at N-3 has been predicted for the ring.⁵⁻⁸ We have now used ^{15}N NMR spectroscopy at $-40\text{ }^{\circ}\text{C}$ to examine the protonation of the aryl pentazoles **1** and **4** (Scheme 1).[†]

Keywords: Pentazole; Low temperature ^{15}N NMR; $-\text{N}_5$ ring; Protonation; Azole.

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[†] Example of the procedure for preparation and purification of the aryl pentazoles: *p*-Anisidine (0.28 g, 2.3 mmol) in MeOH (1.44 cm³) was treated dropwise with concd HCl (0.36 cm³) at $0-2\text{ }^{\circ}\text{C}$. Slowly and with stirring, isoamyl nitrite (0.34 cm³, 2.5 mmol) was added. The solution was allowed to stand for 20 min at $0-2\text{ }^{\circ}\text{C}$ and then diluted with a MeOH–H₂O (1:1 v/v) mixture (4.0 cm³). A layer of petroleum spirit (bp $40-60\text{ }^{\circ}\text{C}$, 22 cm³) was placed over the solution and the mixture cooled to $-35\text{ }^{\circ}\text{C}$. A precooled solution of terminal ^{15}N -labelled sodium azide (0.15 g, 2.3 mmol) in a MeOH–H₂O (60:40 v/v) mixture (1.22 cm³) was injected into the lower aqueous layer. A solid mixture of 1-(*p*-methoxyphenyl)pentazole and *p*-methoxyphenylazide separated and was filtered using a sintered glass funnel with a cooling jacket at $-40\text{ }^{\circ}\text{C}$. The solid was carefully and repeatedly washed with a MeOH–H₂O (60:40 v/v) mixture (1.5 cm³), precooled to $-55\text{ }^{\circ}\text{C}$, to give 1-(*p*-methoxyphenyl)pentazole. This procedure places ^{15}N -labels at the pentazole N-2 and N-3 sites.¹⁴

Protonation of a pyridine-type nitrogen atom in an azole ring causes a shielding shift of up to -100 ppm, in the ^{15}N signal.^{9,10} Using N-2 (N-5) and N-3 (N-4) ^{15}N -labelled *p*-chlorophenyl pentazole **1** and *p*-methoxyphenyl pentazole **4** we have observed ring protonation at N-3. On removal of the proton the neutral pentazole was restored (Fig. 1 and Scheme 1). Figure 1A shows the ^{15}N NMR spectrum of **1** at $-40\text{ }^{\circ}\text{C}$ in $\text{CD}_3\text{OD}-\text{CD}_2\text{Cl}_2$ (1:1 v/v). Figure 1B shows the spectrum of **1** measured in 1:1 (v/v) trifluoroacetic acid–trifluoroacetic anhydride at $-30\text{ }^{\circ}\text{C}$. In this trifluoroacetic acid medium only a low percentage of the pentazole molecules are protonated and the N-3 signal moves upfield by only 11.3 ppm, from 6.5 to -4.8 ppm. This agrees with an exceptionally weak basicity for the ring. However at $-40\text{ }^{\circ}\text{C}$ in the superacid, chlorosulfonic acid ($\approx\text{p}K_{\text{a}} -12$)¹¹ (Fig. 1C, Scheme 1) the N-3 signals of **1** and **4** are shifted upfield by 61.6 and 60.0 ppm confirming full protonation of the pentazole. This is expected for a basic $\text{p}K_{\text{a}}$ of ca. -8.9 . The ^{15}N shift of ca. -61 ppm is the average for the pentazole N-3 and N-4, since N-3 and N-4 are equivalent, and is probably composed of ca. -95 ppm for the protonated N-3 (N-4) and ca. -25 ppm for the adjacent N-4 (N-3) since for tetrazoles and triazoles these are the approximate shifts for protonation at one of the two identical N-atoms, which are β to the substituted pyrrole-type N atom.⁷ When the solution, which gave Figure 1C was carefully treated with $\text{CD}_3\text{OD}-\text{D}_2\text{O}$ (2.2:1, v/v) at $-50\text{ }^{\circ}\text{C}$ HCl gas was evolved generating a dilute H_2SO_4 solution and the ^{15}N NMR spectrum changed to that of Figure 1D. For both of the protonated pentazoles **2** and **5** the neutral pentazole was freed from protonation as confirmed by (a) the restoration of its ^{15}N NMR spectrum (with minor solvent effects), (b) its proton NMR spectrum (Fig. 1D inset), which was identical to that of an authentic sample in the same solvent mixture and (c) separate warming of the solution of regenerated pentazole to $-10\text{ }^{\circ}\text{C}$ whereupon the pentazole slowly decomposed to aryl azide and



Scheme 1. ^{15}N NMR shifts from MeNO_2 for ^{15}N -labelled ($^*\text{N}$) substrates shown with ^1H shifts in parentheses: (i) ClSO_3H ; (ii) $\text{CD}_3\text{OD}-\text{D}_2\text{O}$ (2.2:1, v/v); shifts for starting **1** and **4** in $\text{CD}_3\text{OD}-\text{CD}_2\text{Cl}_2$ (1:1, v/v), and for **1** regenerated from **2** in a mixture of ClSO_3H (1.0 cm^3) and $\text{CD}_3\text{OD}-\text{D}_2\text{O}$ (2.2:1, v/v, 1.0 cm^3).

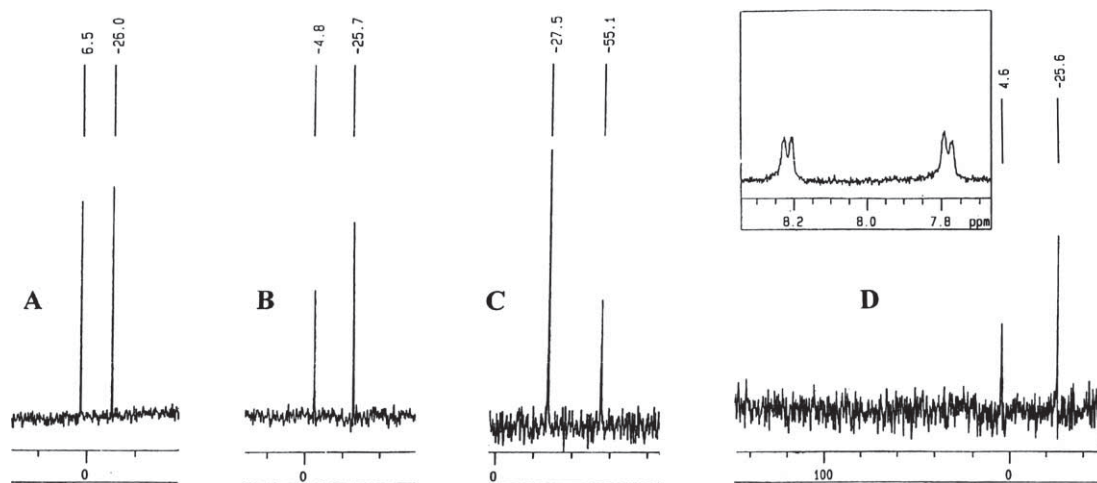


Figure 1. (A, C, D) ^{15}N NMR spectra for **1**, **2** and recovered **1**, respectively, as in scheme; (B) for **1** in $\text{CF}_3\text{CO}_2\text{H}-(\text{CF}_3\text{CO})_2\text{O}$ (1:1, v/v).

N_2 gas (Scheme 1). Hence in chlorosulfonic acid at -40°C 1-aryl pentazoles are protonated at N-3 and do not decompose.

The $-\text{N}_5$ ring is the final member of the azole series, namely a very weak base, which protonates on the N atom, which is β to the pyrrole-type N atom.

It is of interest that **4** protonates on the pentazole N-3 and not on the ethereal oxygen. The $\text{p}K_a$ of protonated anisole is -6.54 .^{11,12} A *p*- NO_2 group lowers the $\text{p}K_a$ of phenol by 2.74 units and since HN_5 is a stronger acid than HNO_3 ¹³ the $-\text{N}_5$ substituent should lower the $\text{p}K_a$ of $\text{MeO}^+\text{H}-\text{C}_6\text{H}_4\text{N}_5$ to ca. -9 . This, combined with the

increased basicity of the pentazole ring due to the electron donating *p*-methoxyphenyl substituent, is enough to direct protonation of compound **4** to N-3.

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